

## Pt-doped RuO<sub>2</sub> electrocatalyst for efficient and durable oxygen evolution in acidic media

**Ranjith Bose**

*Department of Chemical Engineering, Higher Colleges of Technology, Al Dhannah City, Al Dhafra Region, Abu Dhabi, United Arab Emirates.*

[rbose@hct.ac.ae](mailto:rbose@hct.ac.ae)

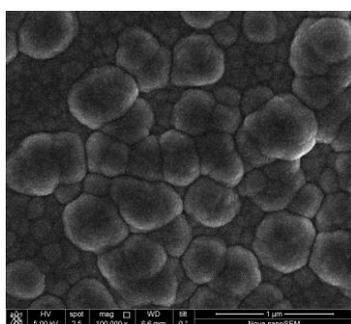
### Abstract

The oxygen evolution reaction (OER) is crucial in energy conversion systems like water electrolysis, CO<sub>2</sub> reduction, and metal-air batteries.<sup>1</sup> It requires high energy to overcome kinetic barriers due to the four-electron transfer reaction.<sup>2-4</sup> In acidic electrolytes, OER offers a higher reaction rate, fewer side reactions, and better cell design, aided by polymer electrolyte membranes (PEM) that enhance proton conductivity and reduce gas crossover. Hydronium ions higher conductivity (350 S cm<sup>2</sup> mol<sup>-1</sup>) than hydroxide ions (198 S cm<sup>2</sup> mol<sup>-1</sup>) enables large current densities, emphasizing the need for cost-effective, durable OER catalysts for acidic media. Notably, Ruthenium oxide (RuO<sub>2</sub>) is a promising OER catalyst in acidic media due to its activity and lower cost than Iridium oxide (IrO<sub>2</sub>); however, it suffers from stability issues.<sup>5</sup> Doping strategies, such as introducing noble metals like platinum (Pt), have shown promise in enhancing RuO<sub>2</sub>'s performance and durability. Considering this, we employed a synergistic approach to develop Pt-incorporated RuO<sub>2</sub>-based catalysts on Titanium (Ti) substrate using a viable electrodeposition technique, achieving simultaneous enhancement of intrinsic activity and stability. The Pt-doped RuO<sub>2</sub>/Ti electrode demonstrates significantly enhanced OER activity in 0.1 M HClO<sub>4</sub> compared to the pristine RuO<sub>2</sub>/Ti electrode, achieving notably lower overpotentials of 194 mV and 206 mV at 10 mA cm<sup>-2</sup>, respectively. Experimental results indicate that incorporating Pt into the RuO<sub>2</sub> lattice increases oxygen vacancy concentration, promoting stronger interactions with reaction intermediates and reducing the energy barrier for OOH\* formation. Furthermore, the enhanced activity is achieved without compromising stability, offering valuable insights for designing high-performance acidic OER catalysts and providing a foundation for practical applications in PEM electrolyzers.

### References

- [1] Xia, C., Jiang, Q., Zhao, C., Hedhili, M. N. & Alshareef, H. N. *Advanced Materials*, 2016, 28, 77-85.
- [2] An, L., Wei, C., Lu, M., Liu, H., Chen, Y., Scherer, G. G., Fisher, A. C., Xi, P., Xu, Z. J. & Yan, C.-H. *Advanced Materials*, 2021, 33, 2006328.
- [3] Zeng, Z., Gan, L. Y., Bin Yang, H., Su, X., Gao, J., Liu, W., Matsumoto, H., Gong, J., Zhang, J., Cai, W., Zhang, Z., Yan, Y., Liu, B. & Chen, P. *Nature Communications*, 2021, 12, 4088.
- [4] Xu, J., Li, J., Lian, Z., Araujo, A., Li, Y., Wei, B., Yu, Z., Bondarchuk, O., Amorim, I., Tileli, V., Li, B. & Liu, L. *ACS Catalysis*, 2021, 11, 3402-3413.
- [5] Chen, D., Liu, T., Wang, P., Zhao, J., Zhang, C., Cheng, R., Li, W., Ji, P., Pu, Z. & Mu, S. *ACS Energy Letters*, 2020, 5, 2909-2915

### Figures



**Figure 1:** SEM image of Pt-doped RuO<sub>2</sub> catalyst.